

Name: _____ Class: _____

Nomenclature Activity

Part A: Binary compounds of metals and nonmetals are called ionic
Highlight these compounds _____

Part B: Binary compounds of nonmetals and nonmetals are called molecular/covalent
Highlight these compounds _____

Part C: Give the formula or the name for each

1. NaCl

Sodium chloride

2. Mercury (II) Chloride

$HgCl_2$

3. HCl

Hydrogen chloride

4. CO_3

Carbon trioxide

5. Potassium cyanide

KCN

6. CaS

Calcium Sulfide

7. Cesium bromide

$CsBr$

8. KOH

Potassium Hydroxide

9. Li_3N

Lithium Nitride

10. Chloride monoxide

ClO

11. CoI_2

Cobalt(II) Iodide

12. ICl

Iodine Monochloride

13. Ag_2S

Silver Sulfide

14. Co_2S_3

Cobalt(III) Sulfide

15. Magnesium fluoride

MgF_2

16. Lithium nitride

Li_3N

17. ~~Co_2F~~ PbO_2

$Pb_2O_4 \rightarrow PbO_2$

lead(IV) oxide

18. $FeBr_3$

Iron(III) bromide

19. NaH

Sodium Hydride

20. Ammonium chloride

NH_4Cl

21. AlI_3

Aluminum Iodide

22. Barium sulfate

$BaSO_4$

+2 +2 Reduce

23. S_4N_4

tetraSulfur tetraNitride

24. CuI

Copper(I) Iodide

25. ZnCl_2

Zinc chloride

26. NaOH

Sodium Hydroxide

27. Sulfur Hexafluoride

SF_6

28. Beryllium Oxide

BeO

29. TiO_2

Titanium oxide

30. Sodium nitrite

NaNO_2

31. Silicon Tetrachloride

SiCl_4

32. Cl_3

Carbon triiodide

33. $\text{Sn}(\text{NO}_3)_4$

Tin (IV) Nitrate

34. $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$

Ammonium dichromate