**Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Endothermic Exothermic Lab Rubric**

\_\_\_\_\_\_/2 Title

\_\_\_\_\_\_/2 Problem

\_\_\_\_\_/2 Hypothesis

\_\_\_\_\_/2 Indicate where heat would appear in the reactions

\_\_\_\_\_/2 Data Table formatted and anchored

\_\_\_\_\_/2 Calculate the temperature change, ∆*t*, for each reaction (Use equation editor) formatted and anchored

\_\_\_\_\_/2 Tell which reaction is exothermic. **Explain.**

\_\_\_\_\_/ 5 Which reaction had a negative ∆*t* value? Is the reaction endothermic or exothermic? **Explain**

\_\_\_\_\_/6 Describe three ways you could tell a chemical reaction was taking place. Use Science terms when possible.

\_\_\_\_\_\_/5 Which reaction took place at a greater rate? **Explain your answer. Please reference your graphs for support.**

\_\_\_\_\_/15 Import your graphs from the Logger Pro software. Label the axis, Final temperature, *t2 with value and unit ,* Initial temperature, *t1 with value and unit, and title the graph. Both graphs should be formatted and anchored MMSTC Style.*

***\_\_\_\_\_\_\_\_\_\_\_/45 pts Total***